

TABLE I

Dye	Solvent	Concentration [M]	Peak Power Out <sup>a</sup> for 5 J Input (KW)	Number of Shots <sup>b</sup> to 50% Decline of Initial Peak Power	$\lambda_{\max}$ , Untuned (nm)
I	ethanol	$7.5 \times 10^{-4}$	2.7	$7 \times 10^3$	460
II	ethanol	$7.5 \times 10^{-4}$	4.1	$5 \times 10^3$	480
III	p-dioxane	$7.5 \times 10^{-4}$	3.8	$3 \times 10^5$	481

<sup>a</sup> Fresh solution.

<sup>b</sup> 500-ml solution.

applications. Most flashlamps would have to be replaced by or before that time.

Since dye (III), 7-diethylamino-4-trifluoromethylcoumarin, is, as far as we know, not commercially available, we give here a brief description of the synthesis procedure we used. The reaction is based on the Pechmann [6] condensation of  $\beta$ -ketonic esters with monohydric phenols.

6 g of *m*-diethylamino phenol and 8 g of ethyltrifluoroacetate were refluxed in 30 ml of ethanol for 15 h with 6 g of anhydrous zinc chloride as a condensing agent. The resultant reaction product was stirred into 1.5 l of water, acidified with 20 ml of concentrated HCl. After 12 h the precipitated solid was collected on a glass-fritted filter and washed with water several times. This crude product was purified by vacuum sublimation and finally recrystallized from absolute methanol. The yield is 49 percent. The material consists of feathery, yellow-green crystals with MP 79–81°C. Analysis: Calculated: C 59 percent, H 4.9 percent, N 4.9 percent, F 20 percent; Found: C 58.9 percent, H 4.8 percent, N 4.9 percent, F 18 percent.

#### REFERENCES

- [1] C. Loth and Y. H. Meyer, "Study of a 1-watt repetitive dye laser," *Appl. Opt.*, vol. 12, pp. 123–125, 1973.
- [2] E. P. Ippen, C. V. Shank, and A. Dienes, "Rapid photobleaching of organic laser dyes in continuously operated devices," *IEEE J. Quantum Electron.* (Corresp.), vol. QE-7, pp. 178–179, Apr. 1971.
- [3] D. Beer and J. Weber, "Photobleaching of organic laser dyes," *Opt. Commun.*, vol. 5, pp. 307–309, 1972.
- [4] P. Sorokin, "Organic lasers," *Sci. Amer.*, vol. 220, pp. 30–40, Feb. 1969.
- [5] K. H. Drexhage, "Design of laser dyes," presented at the 7th Int. Quantum Electronics Conf., Montreal, Que., Canada, May 8, 1972, Paper B.5.
- [6] H.v. Pechmann and M. Schaal, "Studien Über Cumarine II," *Ber.*, vol. 32, pp. 3690–3696, 1899.

#### Correction to "Intracavity Breakdown in CO and CO<sub>2</sub> Lasers"

S. D. ROCKWOOD, G. H. CANAVAN,  
AND W. A. PROCTOR

In the above paper,<sup>1</sup> on page 155, the abscissas of Figs. 1 and 2 are incorrectly labeled. They should read,  $t_g$  (s), and not  $t_p$  as stated. In application to breakdown measurements the conversion from  $t_g$  to  $t_p$  is made by shifting the threshold curves in Figs. 1 and 2 horizontally to the right by the scale factor  $\ln(N/n_0)$  equal to the number of electron exponentiations required to attain breakdown. Since optical attenuation on the order of 4 percent/cm occurs through electron free-free absorption at  $N_e \approx 2 \times 10^{15}/p \cdot \text{cm}^{-3} \text{atm}^{-1}$  for 10.6- $\mu$  radiation and  $N_e \approx 10^{16}/p \cdot \text{cm}^{-3} \text{atm}^{-1}$  for 5.0- $\mu$  radiation, scale factors of 5 and 7 are suggested for CO<sub>2</sub> and CO lasers, respectively. These numbers assume an initial electron density  $n_0 = 10^{18} \text{ cm}^{-3}$  and a gas pressure,  $p = 1\text{-atm}$  at standard temperature. Using these factors the threshold fluence for short pulses becomes  $ft_p \approx 5 \text{ J/cm}^2$  for CO<sub>2</sub> lasers and  $ft_p \approx 28 \text{ J/cm}^2$  for CO lasers. The CW threshold fluxes remain unchanged.

Manuscript received March 7, 1973.

S. D. Rockwood is with the Los Alamos Scientific Laboratory, University of California, Los Alamos, N. Mex. 87544.

G. H. Canavan and W. A. Proctor are with the Air Force Weapons Laboratory, Kirtland AFB, N. Mex. 87117.

<sup>1</sup>S. D. Rockwood, G. H. Canavan, and W. A. Proctor, *IEEE J. Quantum Electron.*, vol. QE-9, pp. 154–157, Jan. 1973.